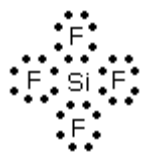
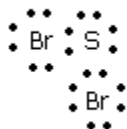


Unit 6 Review Worksheet Answers

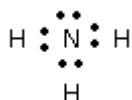
1. less
2. ionic - transfer of e⁻; covalent - share e⁻
3. $> 1.7 =$ ionic; $< 1.7 =$ covalent
 $0 - 0.3 =$ nonpolar covalent; $0.3 - 1.7 =$ polar covalent
4. metal & nonmetal
5. polar covalent - 2 different nonmetals; nonpolar covalent - 2 of the same nonmetal
6. If you have used all of the dots and all atoms don't have 8 dots around them.
7. first element written is central atom (except H)
8. - no unshared e⁻ pairs on central atom
- surrounding atoms must be same element
9. A = central atom; B = surrounding atoms; E = unshared e⁻ pairs on central atom
10. There are 2 surrounding atoms
11. AB₂E & AB₂E₂
12. AB₂
13. AB₄
14. AB₃
15. AB₃E
16. ionic - formula unit; covalent - molecule
17. molecular
18. ionic
19. ionic
20. molecular
21. & 22. hydrogen bonding - b/n molecules of NH₃, H₂O, & HF
dipole-dipole forces - b/n polar molecules
London dispersion - nonpolar molecules & noble gases
23. (D) < (C) < (A) < (B)



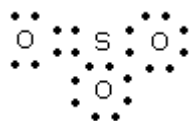
24. (A) AB_4 , tetrahedral, nonpolar, London dispersion



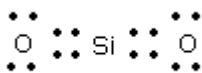
(B) AB_2E_2 , bent, polar, dipole-dipole & London



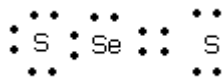
(C) AB_3E , trigonal pyramid, polar, hydrogen bonding, dipole-dipole, London



(D) AB_3 , trigonal planar, nonpolar, London



(E) AB_2 , linear, nonpolar, London



(F) AB_2E , bent, polar, dipole-dipole, London

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